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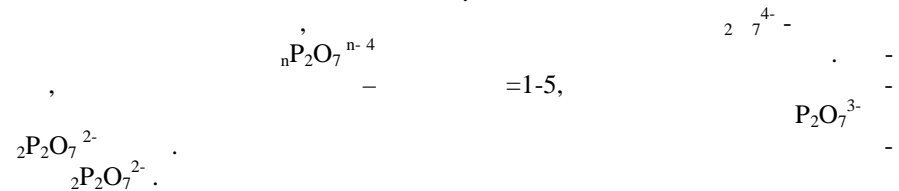
Pyrophosphate palladium complexes electrochemical behavior was investigated using linear step voltammetry. The ligand and central ion concentration effect on the complexes red-ox reactions kinetic parameters were estimated. The chemical reaction of the pyrophosphate palladium complex dissociation exhibits before the charge transfer step.

[1],



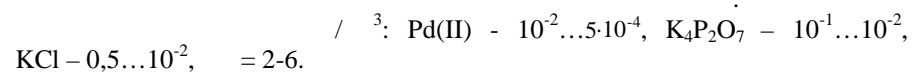
[2,3].

[4],



« . . . ».

« . . . »
 [5].



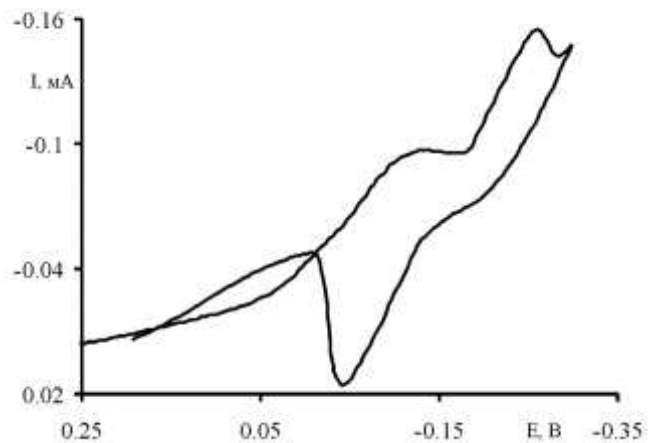
[6],

$$\begin{aligned} & (\quad) \quad (\quad) \\ & s = 0,002 \dots 0,05 / \\ & [7, 8]. \end{aligned}$$

$$1, \quad 5 \leq (\text{P}_2\text{O}_7^{4-}) / (\text{Pd(II)}) \leq 30, \quad n \quad 2.$$

$$\begin{aligned} & (\quad . 1) \quad 0,095 \dots 0,14 \\ & 0,145 \dots 0,16 \end{aligned}$$

(. 1).



1. $\text{Pd(II)} - 10^{-2}$, $\text{K}_4\text{P}_2\text{O}_7 - 10^{-1}$, $\text{KCl} - 0,5$ $s = 0,005$

1

0,05 /

Concentration / 10^{-2}	s	Anodic			Cathodic		
		E_p	$E_{1/2}$	I_p	E_p	$E_{1/2}$	I_p
$\text{Pd(II)} - 10^{-2}$, $\text{K}_4\text{P}_2\text{O}_7 - 10^{-2}$, $\text{KCl} - 0,5$	0,38	-0,27	-0,21	$7,0 \cdot 10^{-5}$	-0,02	-0,085	$7 \cdot 10^{-5}$
$\text{Pd(II)} - 10^{-2}$, $\text{K}_4\text{P}_2\text{O}_7 - 5 \cdot 10^{-2}$, $\text{KCl} - 0,5$	0,34	-0,24	-0,195	$5,7 \cdot 10^{-5}$	-0,03	-0,08	$6,3 \cdot 10^{-5}$
$\text{Pd(II)} - 10^{-2}$, $\text{K}_4\text{P}_2\text{O}_7 - 10^{-1}$, $\text{KCl} - 0,5$	0,3	-0,26	-0,215	$5,7 \cdot 10^{-5}$	-0,04	-0,095	$6,8 \cdot 10^{-5}$

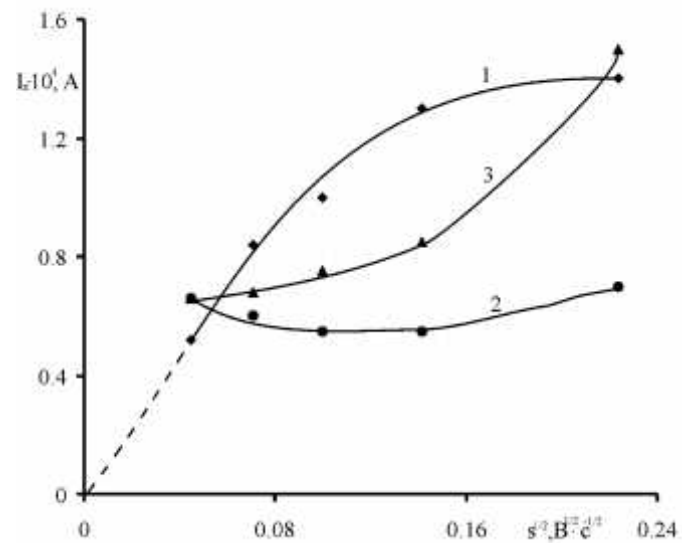
(.2).

$s = 0,23$, $s = 0,28$.

$I \propto s^{-1/2}$

0,12...0,17

(.2),



2. $\text{PdCl}_2 - 10^{-2}$, $\text{K}_4\text{P}_2\text{O}_7 - 10^{-1}$, $\text{KCl} - 0,5$

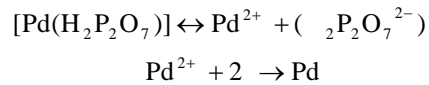
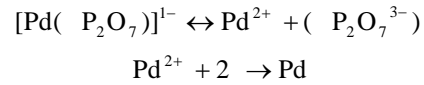
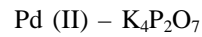
(1), (2), (3)

)

(28,0 , 58,0),
42,0

$$\lg(I/I_d - I) - E$$

$$\lg(I \cdot I_d / I_d - I) -$$



(II)

(II)

/2

(.2).

$\text{Pd}(\text{P}_2\text{O}_7)_2^{6-}$
(/ ³): Pd(II) - 10⁻³, K₄P₂O₇ - 5·10⁻², KCl - 10⁻²

s, /	,	,	/2,	,	,	/2,
0,002	-0,26	4,8·10 ⁻⁵	-0,23	-0,10	4,6·10 ⁻⁵	-0,14
0,005	-0,35	10,0·10 ⁻⁵	-0,27	-0,04	8,2·10 ⁻⁵	-0,09
0,01	-0,41	12,2·10 ⁻⁵	-0,32	0,02	10,0·10 ⁻⁵	-0,04
0,02	-0,49	15,4·10 ⁻⁵	-0,37	0,09	13,4·10 ⁻⁵	-0,02
0,05	-0,69	30,5·10 ⁻⁵	-0,49	0,34	20,5·10 ⁻⁵	0,20

$$s = (\lg I / \lg s)_c,$$

$$s = 0,58,$$

$$s = 0,46.$$

$$s^{-1/2} - s^{1/2}.$$

$$- s^{1/2}$$

$$\lg(I \cdot I_d / I_d - I) -$$

$$\lg(I/I_d - I) - ,$$

85,0).

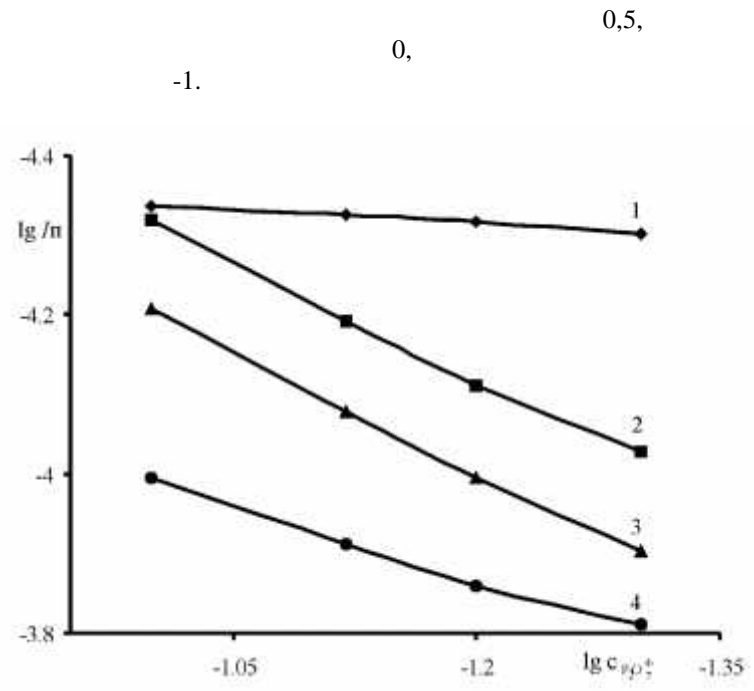
(Pd(II))

- P₂O₇⁴⁻ - (.3).

$$\lg I - \lg c_{\text{Pd(II)}}$$

(s = 0.002 /) 0,

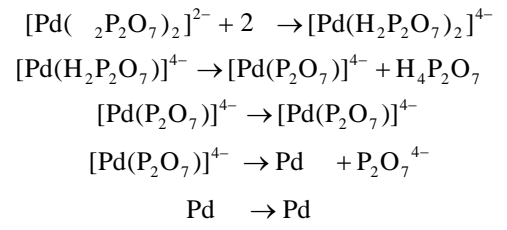
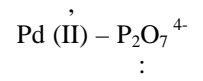
-1.



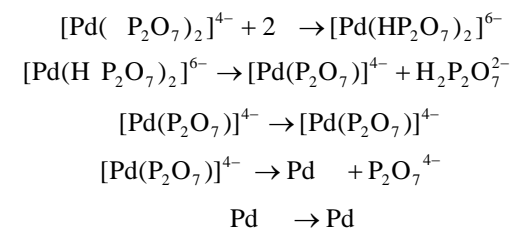
.3.

s=0.002 / (1), s=0.005 / (2), s=0.01 / (3), s=0.02 / (4)

Pd(II)



:



1. 1976.-240 2. 2002. - 3. - 497-502. 3. (II) 1990.- 16.- 8. - 1110-1117. 4. 1980.- 16. - 19. - 1298-1303. 5. 1971.- 310 6. 1991. - 288 7. 2003.-592 8. 1990. - 26. - 1. - 63-68.

10.03.06