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 180 – 200 ° , -
 160 – 180 ° , 80 -
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 140 – 160 ° , 10 – 20 ° . -
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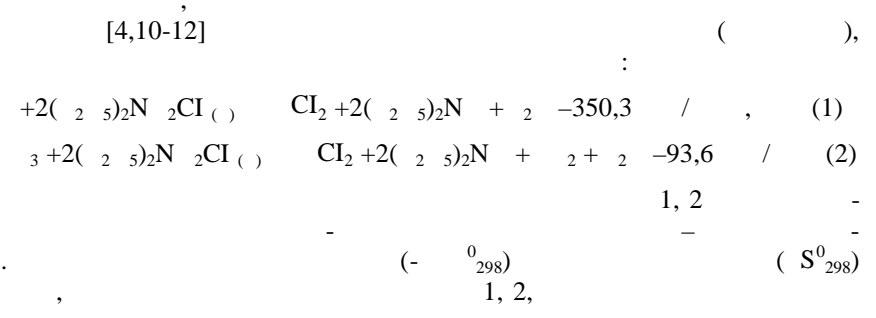
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∴ , ∴ , 296,15 1₂ -
 3. 482,15 -

A thermodynamic possibility found of course of solid-phase interaction reactions between diethylamine chloride, calcium oxide and carbonate. Formation reaction of solid CaCl₂ is shown to take place at a substantial rate already at temperatures above 296,15 K for CaO and above 482,15 K for CaCO₃.

(8 – 10⁻³ /) (220-250 /) [1, 2],
 N₃
 N₃ 99 %
 () [3 – 5].

[6],
 [7-10].



[7, 10, 11, 13].

(S_{298}^0) ($4 \text{ }_{12}\text{NCl}$) -

[9, 10, 12],

S_{298}^0

($4 \text{ }_{12}\text{NCl}$) - 189,98 / [10, 13].

	- S_{298}^0 /	S_{298}^0 /
	- 635,0	39,77
($2 \text{ }_5\text{N}$) $2 \text{ }_{17}\text{Cl}$	- 357,07	189,98
Cl_2	- 797,85	92,09
($2 \text{ }_5\text{N}$)	- 71,16	352,4
2	- 241,95	188,9
3	- 1218,08	92,09
2	- 393,77	213,32

(G^0) :

$$G^0 = S_{298}^0 - S_{298}^0, \quad (3)$$

$$G^0 = -RT \ln, \quad (4)$$

; R - 8,31469 / ;

[8,9]. (2)

Cl_2

1, 2, 23° (296,15)

209° (482,15)

3.

1 2.

1.

G^0 1 2.

	293,15	296,15	323,15	373,15	423,15	473,15	478,15	482,15	523,15	573,15
	$G^0, /$									
($2 \text{ }_5\text{N}$) $2 \text{ }_{17}\text{Cl}$	1632	0	-2,9·10 ⁴	-4,3·10 ⁴	-7,1·10 ⁴	-10·10 ⁴	-	-	-1,5·10 ⁵	-1,6·10 ⁵
($2 \text{ }_5\text{N}$) $2 \text{ }_{17}\text{Cl}$ +	$G^0, /$									
$2^+ 2$	5·10 ¹	1	294,9	1,2·10 ⁶	7,1·10 ⁸	1,1·10 ¹¹	-	-	6,2·10 ¹²	1,8·10 ¹⁴
($2 \text{ }_5\text{N}$) $2 \text{ }_{17}\text{Cl}$	-	-	-	-	-	6,6·10 ³	2,9·10 ³	0	-2,1·10 ⁴	-6,6·10 ⁴
($2 \text{ }_5\text{N}$) $2 \text{ }_{17}\text{Cl}$ +	$G^0, /$									
$2^+ 2$	-	-	-	-	-	2·10 ⁻²	1,8·10 ⁻¹	1,0	6,9·10 ⁶	7,4·10 ³

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