

2. () [4] [5]. [6] [7] [8] [9] : 1) [8]; 2) [10], [11].

[12]. [13-16]. [17] [18] [19] [20] 91 3. () () ; 2) ; 3) 4. : 1) ; 2) ; 3)

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28.69 , " , 700
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 5.19 , 1.34
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 1955 .
 . 83.5-84.5 ,
 - 0.10 .
 0.5 5.0 .
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 (. 81.60-82.80)
 1979 . 2.2 .
 [21],
 : 2004 . - 2.15 , 6-496 (Q)
 - 2.88 1962 . - 1.6 1988 .,
 2004 . 2.01 , - 2.27 .
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6.

: 3.21-3.33
 1 (635 1990).
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 1 2 (. . 1)
 (2) 0.104 / 7
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 , / : 1) 0.124; 2)
 0.107; 3) 0.017,
 0.013; 4)
 1990 . 0.111. ()
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 94.3% , ()
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	Pb, /					
	26.07. 1999	10.07. 2000	8.08. 2001	29.05. 2002	19.11. 2003	29.04. 2004

-	1.	-	37	18	62	28	73	79
	2.	-	72	68	62	35	85	74
	3.	-	52	51	300	686	100	17
	4.	-	1	3	2	14	8.7	13
	5.	-	2	9	2	11	9.3	13
		-	-	-	-	*) =±25	**) =±40	
: *) [22] [8], .21-22, .26-27; **) [23].								

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1, 2 (50 66 /).
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2003 . 3
686 100 / . 2

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1	18	25	34.5
2	24	56	50.3

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[7] (-)
4472-84). [8] 26640-85 ()
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[6]. - - - - -

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8.

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; 2) -

17.4.3.01-83, 17.4.4.02-84. -

[22]. -

[23] 2004 ., [24] 2003 . -

(), [25]. -

1, 2 -

9. () -

CG1-CG6. -

CG1. (CG1.1) (CG1.2) , -

CG2. (-

CG2.1 -

CG2.2 , () . -

CG3. -

(CG3.1), (CG3.2) (CG3.3) -

>C), (C > >), (-

(> >) . -

CG4. (CG4.1), -

97 / -

CG5. (CG5.1) (CG5.2). -

=32 / 42-128-4433-87, -

[26]. -

CG6. () . -

10. " " -

1. 1, 2 , -

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CG1.1-CG1.2

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CG2.1.

CG2.2

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1, 2

CG3

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4.

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1, 2

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5.

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17.4.2.01-81

32 /

1, 2.

CG5.2).

[27,28],

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CG5.2.

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[29].

CG6

: **CG1.1,CG2.2,CG3.3,CG5.1,CG6.**

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19 1996 . N 173. 6.

" , . 20. 7. " " . 8.

" , . 1,7. 9. " . 10. -89-80

. 11. . 2.2.-1-2003

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. 12.

(. 27.10.97 . N 171,
5 1998 . N 285/2725). 13.

, 2001.- 244 . 14.

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11.

(CG1,CG2,CG3,CG4,CG5,CG6)

101

, 2002, . 4.- . 163-169. 16.

)// , 2002, . 4.- . 111-117. 17.

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, 1991, . 75-83. 18.

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" " " " - 2003 - 3 - . 121-126. 19.

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, 1991, . 95-102. 20.

// : . . / . , 1993, . 55-62. 21.

2004 /

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03.11.2003 N 98. 23.

081/12-0009-01.

. 24.

16.01.89. 25.

42-128-4433-87, , 30.10.97 4433-87. 26.

2004, 2, . 10-12. 27. // , 1986.- 320 . 28. / .- .: 1976. 29. () / " -2005". , . 24-27 2005 .- .: " , 2005.- . 518-524. 18.05.05

621.357.12

PbO₂

PbO₂

The explored process electroplating of PbO₂ from two alkaline electrolytes - an plumbic alkalinitu and electrolyte on base EDTA. It is determined diffusing ability both electrolyte. Advantage of the electrolyte are Shown on base EDTA. The Proved possibility of the reception bipolar electrode with plumbic dioxide by anode side and stood by anode net.

[1].

[3].

[3].

[4-6].

[7].

: 1) Pb²⁺ - 0,2 - / ³ (Na₂PbO₂), NaOH