

614.84

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.

Thermodynamical computations cooling operation line carbonates and ammonium salts were carry out. Adjusted, that fluoride and ammonium sulphate has greater, than water cooling action. To increase cooling action water suggests used dopers for water ammonium sulphate. These solutions have higher extinguishing operation then water.

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().

(2254 / 25 °).

[1]

[2]

25 – 600 ° .

(20 25 °) 100 ° .

– 20 °

(),

1000 ° .

500 – 800° .

[3 – 4].

[3 – 4].

230 – 260 ° [5].

230 – 260 ° .

(100 °)

[2]

(NH₄)₂CO₃, NH₄HCO₃.

: MgCO₃, MnCO₃, FeCO₃,

[6].

°₂₉₈

1.

1

(°₂₉₈)

	MgCO ₃	MnCO ₃	FeCO ₃	(NH ₄) ₂ CO ₃	NH ₄ HCO ₃
° ₂₉₈ , /	1202,4	896,5	685,3	1782,3	1572,6

(101,3).

[7]

100 °

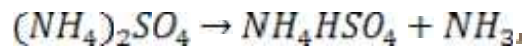
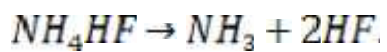
2

25 °

2

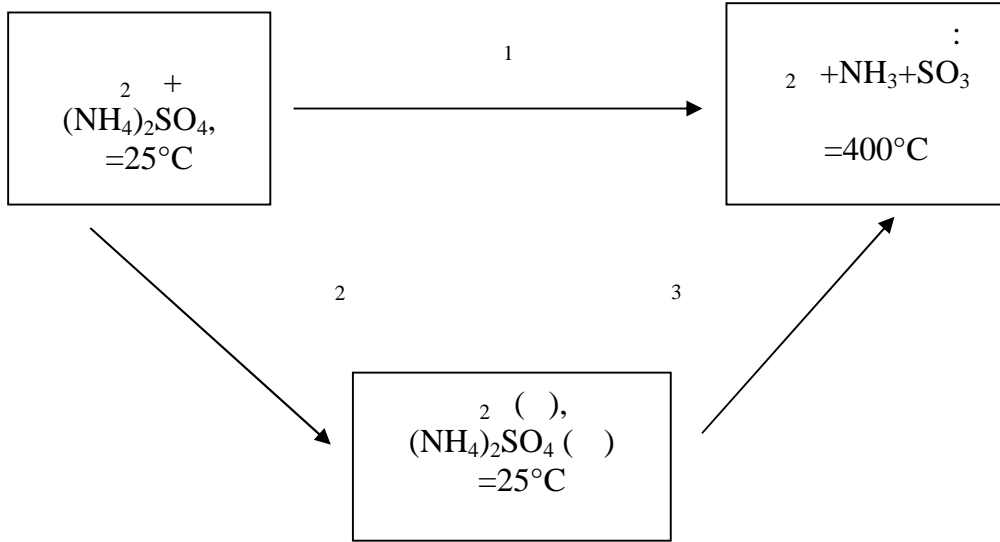
(), (°₂₉₈)
(S)

	NH ₄ F	NH ₄ Cl	NH ₄ Br	NH ₄ I	(NH ₄) ₂ SO ₄
, °	238	338	394	405	>218
° ₂₉₈ , /	3895	3284	1939	1250	3815
S, /100	37,0	37,2	74,2	172,3	76,9



(, (VI)).

25 ° 1
 1 30 %
 $(\text{NH}_4)_2\text{SO}_4$.
 [1].
).



25°C
 400°C

$$I = 2 + 3 \quad (1)$$

$(\text{NH}_4)_2\text{SO}_4$, 1
 $(\text{NH}_4)_2\text{SO}_4$, 3 -
 $(\text{NH}_4)_2\text{SO}_4$, 2 -
 :

$$2 = - \quad (2)$$

(2) (1) :

$$I = 3 - \quad (3)$$

m_3 (m_2),
 ($m_3((NH_4)_2SO_4)$). m_1
 [1] (m_2) = 3324 / ($m_3((NH_4)_2SO_4)$)
 : 25 400 ° ,
 , 0,3
 :

$$m_3((NH_4)_2SO_4) = m \cdot C_p((NH_4)_2SO_4) \cdot T + m \cdot C_p = 0,3 \cdot 1,42 \cdot (400 - 25) + 0,3 \cdot 3815 = 1305$$

0,7 .

$$m_2 = 3324 \cdot 0,7 = 2327$$

$$m_3 = m_3((NH_4)_2SO_4) + m_2 = 1305 + 2327 = 3632 / ()$$

[8, 9].

1 30 % $(NH_4)_2SO_4$ 159 /

$$m_1 = m_3 - 159 = 3632 - 159 = 3473 / ()$$

5 % , .

