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–   . 127 – 137. **2.**   . .   ,   . .   .

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The comparative analysis of existent electrolytes of nickelage and silvering which composition of solution and modes of electrolysis for the receipt of coverages with the set functional properties (atalytic activity to process of oxidation carbon oxide) is offered on the basis of is resulted in the article. Analysis of composition of existent electrolytes from which besieging of nickel with the small quantity of silver testifies possibly, that the best solutions be sulfamate, which are widely used for forming of coverages of nickel of a various thickness.

-

$$1 - 5\ \%$$

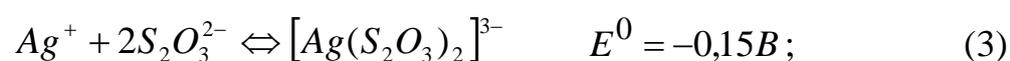,   .

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,   [1].

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[2].

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Ni – Ag   -

(20 %   )   ,

50   .

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,   ,   -

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,   -

.   .

$$(\quad 100\quad)$$

$$-$$ -

$$(1 - 5\ \%\qquad\qquad)$$

-   .

[3],

$$E^0_{Ni^{2+}/Ni} = -0,25 B\,,\qquad\qquad E^0_{Ag^+/Ag} = +0,799 B\,.$$

-

1,04   -

.   ,

,

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,   ,   ,   ,

,   [4, 5]:

$$Ag^+ + 2CN^- \Leftrightarrow \left[Ag(CN)_2\right]^- \qquad E^0 = -0,31 B; \qquad (1)$$

$$Ag^+ + 3J^- \Leftrightarrow \left[AgJ_3\right]^{2-} \qquad E^0 = 0,017 B; \qquad (2)$$

$$Ag^+ + 2S_2O_3^{2-} \Leftrightarrow \left[Ag(S_2O_3)_2\right]^{3-} \qquad E^0 = -0,15 B; \qquad (3)$$

$$Ag^+ + CNS \Leftrightarrow \left[Ag(CNS)\right] \qquad E^0 = 0,09 B; \qquad (4)$$

$$Ag^+ + 2SO_3^{2-} \Leftrightarrow \left[Ag(SO_3)_2\right]^{3-} \qquad E^0 = 0,295 B; \qquad (5)$$

$$Ag^+ + P_2O_7^{4-} \Leftrightarrow \left[Ag(P_2O_7)\right] \qquad E^0 = 0,15 B; \qquad (6)$$

$$Ag^+ + 2NH_3 \Leftrightarrow \left[Ag(NH_3)_2\right]^+ \qquad E^0 = 0,373 B. \qquad (7)$$

(1 – 7),

-

(NH$_2$SO$_2$OH),   (   )   ,

-

Ag^+ -

0,12 – 0,3   / ^{3},   0,2 – 0,7   / ^{3}.

,   -

,   ,

NH$_2$SO$_2$OH.

0,01 – 0,25   / ^{3},   0,3   / ^{3}.   -

.

-

$= 2,5 \cdot 10^{-14}$

$$E^0_{Ag^+/[Ag(S_2O_3)_2]^{3-}} = 0,017 B$$

0,773

		Ag$^+$, / ^{3}		
		0,5 / ^{3} NH$_2$SO$_2$OH	1 / ^{3} NH$_2$SO$_2$OH	0,64 / ^{3} Ni(NH$_2$SO$_3$)$_2$
Ag(NH$_2$SO$_3$)	–	~ 0,12	~ 0,2	0,7
AgCl	2,04 10^{-3}	. .	. .	. .
AgCl + (NH$_4$)$_2$SO$_4$	9,31 10^{-8}	. .	. .	. .
AgCl + K$_4$P$_2$O$_7$	10^{-4}	. .	. .	. .
AgCl + KJ	1,4·10^{-14}		–	
Ag$^+$ Na$_2$S$_2$O$_3$	4,5·10^{-8}	. .	. .	
AgCl + Na$_2$S$_2$O$_3$	2,5·10^{-14}	~ 1·10^{-2}	2,5·10^{-2}	0,3
Ag	–	0,2	0,3	0,2

Ni – Ag

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    . .   . –   :   , – 1968 . – 1462 . 3.   -
    . .,   . .,   . .   . – .:   , 1987.
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