The main parameters of chemical reactions are defined in article on as-new experimental data, required for modeling of process a thermophotocatalytic transformation. As well as is analysed about-cession an decomposition of pesticides of preparation DDT thermophotocatalytic method.

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[2].

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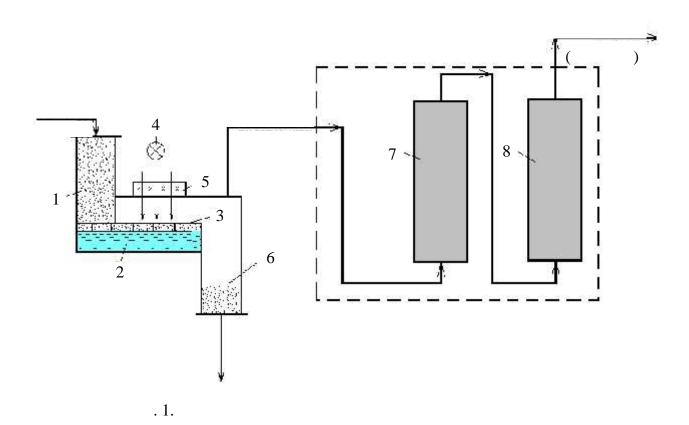
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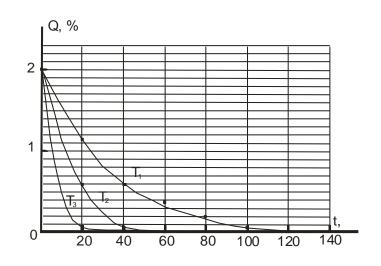
, 5, 3, 1 6, 7, 8 . 1. 1. 3 – 5 3 200-250 ° . 2, 5 20 – 30 (). 6 8 (), 30°, ($_1$ = 200 $^{\circ}$, $_2$ = 250 $^{\circ}$ $_3 = 300 \, ^{\circ} \,)$ (. 2). $k = \frac{2.3}{t} \lg \frac{Q_1}{Q_2}$ (1) ; t k – , ();

 Q_1 Q_2 - t (/).

, $_{1}=200^{0} \qquad \qquad :$

$$\frac{2.3}{t_1} \lg \frac{Q_1}{Q_2} = \frac{2.3}{t_2} \lg \frac{Q_1}{Q_3},\tag{2}$$

 t_1 t_2 - ; Q_1 - - , (/); Q_2 Q_3 - (/) t_1 t_2



. 2.

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$$\frac{2.3}{20} \lg \frac{20}{8.6} = 0.04215 \quad \neq \quad \frac{2.3}{40} \lg \frac{20}{14} = 0.00669. \tag{3}$$

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$$_{1} = 200^{0}$$

‡ = 26 ;

$$_2 = 250^{0}$$

$$\ddagger = 14.5$$
 ; $= 300^{\circ}$

$$= 300^{\circ}$$

:

$$k = -\frac{\frac{a}{2}}{\ddagger_{1/2} a \left(a - \frac{a}{2}\right)} = -\frac{1}{\ddagger_{1/2} a}$$
 (4)

$$k^{200} = -\frac{1}{26 * 20} = -\frac{1}{520}$$

$$k^{250} = -\frac{1}{14.5 * 20} = -\frac{1}{290}$$

$$k^{300} = -\frac{1}{10*20} = -\frac{1}{200}$$

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$$2.3\lg\frac{k_2}{k_1} = \frac{E}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right). \tag{5}$$

$$E = \frac{2.3 R \lg \frac{k_2}{k_1}}{\left(\frac{1}{T_1} - \frac{1}{T_2}\right)}.$$
 (6)

31.10.2007

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It was researched the chemical method of immobilization of ferments with a cross-linking agent – an glutaric aldehyde, as carrier was used silica gel of the mark KSK-G, ferment – amylase. The optimum parameters of the undertaking covalent immobilization were determined.

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